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25 µL of the hybridoma supernatant into rows A1–A12, and is serially diluted 2-fold down the plate. The plate was allowed to incubate for 2 h at 37 °C, which was then followed by a deionized water wash/shaking (10 times). All wells were next treated with 25 µL of a goat anti-mouse glucose oxidase (Cappel) solution, which was diluted previously 1:500 in Blotto. The plate was allowed to incubate for 1 h at 37 °C. This was followed by (10X) deionized water wash/shaking. The plate is finally developed with 50 µL/well of a developing solution (25 mL, 0.2 M sodium phosphate, pH 6.0, 3 mL 20% glucose/deionized water, 200 µL 0.1% horseradish peroxidase/0.1 M phosphate buffer, pH 6.0, 200 µL ABTS (BMB) (45 mg/mL of deionized water). After 30 min the plate (each well) absorbance was determined at 405 nm on a Molecular Devices (V-max) kinetic plate reader.

A typical competition ELISA was performed in the following manner. A "limiting" antibody and antigen concentration as determined above is used in the following manner. The BSA–hapten conjugate (concentration determined above) in PBS is added at 25 µL/well to an ELISA plate. Methanol fix was added at 50 µL/well for 5 min. The methanol is removed, and the plate is air-dried for 10 min. The plate is blocked as described above with Blotto (50 µL/well) at 37 °C for 30 min, and the excess Blotto is removed. Next, the antigen (hapten) or substrate (diluted 2-fold down the plate) is added (described above). After incubation for 1 h at 37 °C, the plate is again washed with deionized water (10X) followed by the addition of developing solution (described above). Absorbances are read at 405 nm.

The data (absorbances) obtained were fitted to eq 2 where CP is the absorbance determined and t is the antigen or substrate concentration.

\[
CP = \frac{P(1)}{1 + P(2)c} + P(3)
\]

**Coupled Kinetic Analysis of Cleavage of DNA by Esperamicin and Calicheamicin**

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Abstract: A coupled kinetic analysis of esperamicin, calicheamicin, and DNase I cleavage of covalently closed circular PM2 DNA has been carried out. Analysis of the optical density data derived from agarose gel electrophoresis experiments shows that esperamicin A₄, like the hydrolytic enzyme DNase I, produces mainly single-strand breaks in DNA. These agents cause covalently closed circular form I DNA to be initially converted to nicked circular form II DNA. However, the ratio of the rate constant for this process \(k'(1)\) to that associated with conversion of form II to linear form III DNA \(k'(2)\) is not consistent with completely random nicking, and some double-strand cleavage may occur. The values of \(k'(1)/k'(2)\) found for DNase I and esperamicin A₄ were 5.4 and 3.9, respectively. The behavior of these agents sharply contrasts with that of esperamicin C and calicheamicin, for which double-strand cleavage of DNA is deduced from the analysis. Although the ratio constant for introducing the first break in DNA for calicheamicin is lower than the corresponding rate constant for esperamicin C, the second break (in the opposing strand) is fast for calicheamicin, making it the better double-strand cleaving agent. These drugs are unique among antitumor agents in that a single activation event on the warhead portion produces a double-strand break in DNA without the need to posttreat the DNA with other agents in order to induce a cleavage. The coupling kinetics are discussed in terms of the structural differences in these unusual anticancer drugs.

The newly discovered anticancer drugs esperamicin¹,² and calicheamicin³,⁴ (Figure 1) exhibit high potency against murine tumor lines. These compounds possess sugar groups appended to an unusual 1,5-diyn-3-ene core structure referred to as the "warhead". In the presence of reducing agents and DNA the warhead undergoes a rearrangement to create a phenylene di-radical that causes strand breaks by attacking the sugar groups.


DNA Cleavage

![Diagram of DNA Cleavage]

An interesting feature of esperamicin and calicheamicin is that a single drug molecule, after activation, can react with opposing strands to produce a double-strand break in DNA. This raises the possibility that a single drug molecule, after activation, can react with opposing strands to produce a double-strand break in DNA. Since this type of lesion is more difficult to repair than a single-strand peak, it is possible that a high frequency of double-strand breaks in DNA may be part of the basis for the extreme cytotoxicity of this unusual class of antitumor agent.

The evidence that esperamicin and calicheamicin cleave DNA via a double-strand process is primarily based on the drug-induced cleavage pattern on opposite DNA strands. In the case of calicheamicin on labeling experiments showing that both radicals (1) and (2) are shown on the right of the figure. If two breaks on opposing DNA strands were relatively close to one another, the DNA segment between the breaks would "melt", causing a break in the opposite strand. In the case of neocarzinostatin, a "double-strand" break is the result of a break in the sugar-phosphate backbone on one strand and the creation of an apurinic/apyrimidic site two nucleotides away from the initial break on the opposing DNA strand. Treatment of the modified DNA with endonuclease IV or pareticine, agents that attack apurinic/apyrimidic sites, is necessary in order to cause the second break in DNA.

In this paper we examine the ability of esperamicins A1 and C and calicheamicin γ1 to cleave closed circular PM2 DNA. The cleavage data are quantitatively analyzed with use of coupled kinetic models representing single- and double-strand cleavage processes. For comparative purposes, we also analyze cleavage of PM2 DNA by the hydrolytic enzyme DNase I under conditions previously reported to yield single-strand breaks in DNA.

PM2 DNA, a covalently closed circular DNA molecule that is 9300 nucleotides long, is a convenient substrate for studying the cleavage action of DNA-damaging enzymes. An agent capable of binding to DNA and yielding 1 equiv of a DNA-damaging group from a single activation event would cause an initial break in the PM2 genome, leading to the loss of supercoiling and the formation of nicked circular form II DNA. Subsequent binding and activation events would cause additional nicking of form II DNA. If two breaks on opposing DNA strands were relatively close to one another, the DNA segment between the breaks would "melt", causing form II to be converted to the linear form, form III DNA. Although DNA melting depends strongly on sequence as well as other factors, breaks on opposing strands up to 6-10 base pairs apart would likely cause linearization of the DNA molecule under the conditions used in this study.

**Experimental Section**

The esperamicins (esp) were obtained from J. Golik of Bristol-Myers Squibb Co., Wallingford, CT, while the sample of calicheamicin γ1 was obtained from G. Elledstad of Lederle Laboratories. Stock drug concentrations (~3 × 10⁻⁴ M) were made by dissolving weighed amounts of the drugs, 1-3 mg, in ethanol. Dilutions necessary for the strand scission work were done by adding a portion of the ethanol stock to a larger volume of the aqueous buffer used in the study. The maximum concentration of ethanol introduced by this procedure in any of the cleavage experiments was <1%. All DNA concentrations are stated in base pairs.

**Cleavage of PM2 DNA.** The cleavage of PM2 DNA by esperamicin and calicheamicin was carried out in a total volume of 13 μL at 37 °C for 10 min in the buffer 6.2 mM Tris-HCl/0.6 mM EDTA, pH 7.5. To 2 μL of the solution containing DNA was added 8 μL of a solution containing the drug, followed by 3 μL of the solution containing dithiorthiol (DTT). The final concentrations of PM2 DNA and DTT present in the reaction medium were 9.4 × 10⁻⁹ M and 0.45 mM, respectively. Concentrations of esp A1: 0, 9.23 × 10⁻¹⁰ M; 1.38, 6.62, 9.23 × 10⁻⁹ M; 1.0; 9.23 × 10⁻⁸ M; 5.0; 9.23 × 10⁻⁷ M; 10.0 μg/mL.

![Diagram of DNA Cleavage]

**Figure 1.** Structures of esperamicins A₁ and C and calicheamicin γ₁ of DNA. An interesting feature of esperamicin and calicheamicin as well as the chromophore of neocarzinostatin is that a single activation event produces 2 equiv of a species capable of reacting with sites on DNA, i.e., the phenylene diradical. This raises the possibility that a single drug molecule, after activation, can react with opposing strands to produce a double-strand break in DNA. Since this type of lesion is more difficult to repair than a single-strand peak, it is possible that a high frequency of double-strand breaks in DNA may be part of the basis for the extreme cytotoxicity of this unusual class of antitumor agent.

**Figure 2.** Photograph of agarose gel showing esperamicin C cleavage of PM2 DNA. The drug concentration in the various lanes, in order of increasing drug concentration (left to right), can be found in the Experimental Section. Lane 1 is zero drug concentration. The various forms of PM2 DNA, covalently closed circular (I), nicked circular (II), and linear (III), are shown on the right of the figure.
The intensities for forms I, II, and III DNA resulting from DNase I cleavage of PM2 DNA as a function of DNase I concentration. Intensities calculated by the single-strand (A) and double-strand (B) models described in the text are open symbols. Curves are fits to calculated intensities.

1.39, 1.85, 4.62, 9.23 × 10^4 M; 1.39, 1.87, 4.62 × 10^3 M. Concentrations of esp C: 0, 9.23 × 10^4 M; 1.85, 4.62, 9.23 × 10^3 M; 1.39, 1.85, 2.31, 4.62, 9.23 × 10^3 M. Concentrations of calicheamicin used in the study: 0, 9.69 × 10^3 M; 1.38, 4.62, 9.23 × 10^3 M; 1.85, 4.62, 9.23 × 10^3 M; 1.38, 1.85 × 10^4 M. After a digest time of 10 min, the reaction was terminated by the addition of 4 μL of a 3 × 10^4 M sodium dodecyl sulfate "stop" solution, and the DNA was loaded onto a horizontal 0.8% agarose gel and electrophoresed. After electrophoresis, the gel was soaked in a 0.5 μg/mL solution of ethidium bromide for 1 h, placed on a UV transilluminator, and photographed with use of a Polaroid camera equipped with positive/negative Type 55 film. The resulting negative was "fixed" according to the manufacturer's recommendation with aqueous 18% Na2SO4. A photograph of the stained gel for esp C is shown in Figure 2.

Cleavage of PM2 DNA was also carried out with use of DNase I in a total volume of 7 μL of the buffer 14.3 mM Tris-HCl/9.1 mM MgCl2/2.3 mM CaCl2, pH 7.5. After proceeding for 10 min, the reaction was terminated by addition of aqueous EDTA and the products were loaded onto an agarose gel. Concentrations of DNase I (units/μL) present in the various reactions: 1.1, 1.67, 2.22, 5.56 × 10^4, 1.1, 1.67, 2.22, 5.56 × 10^3, 1.1, 1.67, 2.22, 5.56 × 10^2, 1.1, 1.67, 2.22, 5.56 × 10^1. 1.1, 1.67, 2.22, 5.56 × 10^-1.

A Molecular Dynamics Model 300A microdensitometer capable of whole-spot integration was used to scan the photographic negatives of the gels. Integration was terminated by addition of aqueous EDTA and the products were loaded onto an agarose gel. Concentrations of DNase I (units/μL) present in the various reactions: 1.1, 1.67, 2.22, 5.56 × 10^4, 1.1, 1.67, 2.22, 5.56 × 10^3, 1.1, 1.67, 2.22, 5.56 × 10^2, 1.1, 1.67, 2.22, 5.56 × 10^1, 1.1, 1.67, 2.22, 5.56 × 10^-1.

The resulting negative was "fixed" according to the manufacturer's recommendation with aqueous 18% Na2SO4. A photograph of the stained gel for esp C is shown in Figure 2. The band intensities are proportional to the amounts of the various forms of DNA present at the end of the digest period τ (10 min). The kinetic models for the interconversion of the various forms of DNA are discussed in the next section. Given such a model, one can calculate the amount of each DNA form for each concentration of cleavage agent, in terms of the rate constants and the amounts of the three forms present before reaction with cleavage agent. The rate constants and initial amounts are parameters determined by seeking the values that minimize the deviation between the experimental and calculated band intensities. Thus, if \( I_{ij} \) is the intensity of form \( i \) at \( t = \tau \) when the cleavage agent concentration is \( c_c \) and if \( I_{ij}^{\text{calc}} \) is the corresponding value calculated according to a model, we minimize

\[
D = \sum_i \sum_j (I_{ij} - I_{ij}^{\text{calc}})^2
\]


The calculated values of $D$ indicate in what way the parameters should be changed to obtain lower values of $D$, and the process is repeated.

Results and Discussion

In the cases of esperamicin and calicheamicin, an activation agent has been postulated to produce a diradical possessing two DNA-damaging equivalents. If both radicals on the phenylene moiety reach DNA targets, strand scission will be a two-step process with reaction of one radical with a site on DNA, e.g., removing a hydrogen atom from deoxyribose, followed by the reaction of the second radical with another DNA site. The likelihood of a concerted process, i.e., identical reaction rates for the reaction of the two radicals, seems low since it would require optimal positioning of both radicals relative to their respective DNA targets. If the drug did not move from its original position in the time required for both radicals of the phenylene moiety to react with DNA targets on opposing strands, the covalently closed circular form I DNA would be converted directly to form III DNA without passing through the nicked circular form II DNA. Although the rate constant of hydrogen atom abstraction of the phenylene diradical of esperamicin and calicheamicin from a donor has not been measured, studies on related systems suggest that this constant is large, $\sim 10^6$ M$^{-1}$ s$^{-1}$, and thus the drug may not have time to migrate to a new position on DNA before hydrogen abstraction from some source occurs. If only one of the phenylene diradicals reacts with DNA or both react with the same strand of the polymer, form I would convert to form II DNA.

Extensive random cleavage of the latter form would eventually produce the linear product, form III DNA.

The experimental results of the studies involving PM2 DNA are shown in Figures 2–6, solid symbols. In the cases of DNAse I and esp A$_1$ (Figures 3 and 4), cleavage caused closed circular form I DNA to decrease in amount while at the same time nicked circular form II increased in amount. Higher concentrations of DNAse I or esp A led to a decrease in the amount of form II while linear form III DNA increased in amount. This behavior contrasted with that of esp C and calicheamicin, which showed decreases in the amount of form I with concomitant increases in the amount of form III DNA (Figures 5 and 6). These intensity/concentration plots (Figures 3–6) were analyzed according to the single- and double-strand cleavage models given below. The calculated rate constants can be found in Table I.

To model the interconversion of the three forms of PM2 DNA, we begin by assuming that the rate of nicking of each form of DNA is proportional to its concentration and to the concentration of the cleavage agent (DNAse I or drug). Each interconversion rate is proportional to the rate of nicking. Then, for example, the rate of decrease of the concentration of the closed circular form I DNA is $d[I]/dt = -k[I][P]$ where $k'$ is a second-order rate constant and $[P]$ is the concentration of cleavage agent. The experimental data for esperamicin and calicheamicin show that there is little or no change in the concentrations of the three forms of PM2 DNA over the digest time unless the drug concentrations exceed about 0.02 $\mu$M. Since the concentration of PM2 DNA (plasmid per unit volume) is only 0.01 $\mu$M, and since we believe almost all the drug is bound to the DNA, this implies that only a fraction of the drug is activated, so that the drug concentration...
The linear form can be produced from the nicked form, or directly to the linear form, \( \text{form I} \) to \( \text{form II} \) to \( \text{form III} \) DNA. Model ds is for double-strand cleavage, form I \( \rightarrow \) form III and form III \( \rightarrow \) form I. 

Second-order rate constants in \( \text{mM} \cdot \text{min}^{-1} \). The quantity \( D \) is the sum of the squared deviations between experimental and calculated data. Stated errors are changes producing 10% increase in \( D \). For DNAse I, the rate constants are in (units/L) \( \cdot \text{min}^{-1} \). There are, as in the previous model, six parameters: three rate constants and three initial concentrations. 

**Derivation of Rate Constants.** For either of the two models, the amounts or intensities of all three forms may be calculated for all drug or DNAse I concentrations in terms of six parameters: three rate constants and three initial concentrations \( (C_0, N_0, L_0) \). For any set of values for these parameters, calculated intensities may be compared with the experimental intensities. The parameters are varied to minimize the sum of the squared deviations: 

\[
D = \sum (I_i - I_i^{\text{cal}})^2 
\]

Here, \( I_i \) is the intensity of form \( i \) when the drug or DNAse I concentration is \( c \). The values of the second-order rate constants found by minimization of \( D \), for each model and for each cleavage agent, are given in Table I, along with the minimum values of \( D \). In Figures 3–6, we show the intensities calculated according to the models, using the best values for the parameters (open symbols). 

Comparison of the values of \( D \) for different models shows whether one model is significantly better than another, or whether the two fit the experimental data equally well. How well the experimental data are fit may also be judged from Figures 3–6 themselves. Comparison of rate constants helps in understanding the physical meaning of the models. Due to the fact that the DNAse I concentration is expressed in units and the drug concentration in moles per liter, the rate constants for the enzyme and drug are not directly comparable. However, comparing the ratios of rate constants is valid. It should be noted that the value of \( k_i' \), the rate constant for conversion of form III DNA to shorter fragments, cannot be determined accurately from the data, because the data are for low drug concentrations, for which the amount of short fragments produced is small.

**DNA Cleavage Mechanism.** DNAse I. Prior to analyzing the cleavage data for esperamycin and calicheamicin using the outlined kinetic approach, we analyzed cleavage of PM2 DNA by the enzyme DNAse I. This enzyme binds and cleaves DNA by a well-known mechanism and, under the conditions used in the study, should cleave via a single-strand process. 

DNAse I is a 30-kDa protein that hydrolyzes the phosphodiester backbone of DNA. Single-crystal X-ray analysis has shown that, when DNAse I is bound to DNA, the calcium ion at the catalytic site of the enzyme is positioned near a phosphate group on one strand of DNA. 

The metal ion acts as a Lewis acid and, when bound to the phosphate group, facilitates the hydrolysis of the diester linkage of DNA. 

**Table I. Rate Constants for Esperamycin, Calicheamicin, and DNAse I Cleavage of PM2 DNA**

<table>
<thead>
<tr>
<th>cleavage agent</th>
<th>model</th>
<th>( k_i' )</th>
<th>( k_4' )</th>
<th>( k_2' )</th>
<th>( D \times 10^5 )</th>
</tr>
</thead>
<tbody>
<tr>
<td>DNAse I</td>
<td>ds</td>
<td>0.49 ± 0.07</td>
<td>0.46 ± 0.20</td>
<td>0.09 ± 0.03</td>
<td>0.11 ± 0.03</td>
</tr>
<tr>
<td>DNAse I</td>
<td>ss</td>
<td>1.95 ± 0.25</td>
<td>1.31 ± 0.21</td>
<td>0.05 ± 0.03</td>
<td>0.10 ± 0.05</td>
</tr>
<tr>
<td>esperamicin</td>
<td>ss</td>
<td>0.31 ± 0.03</td>
<td>0.31 ± 0.03</td>
<td>0.50 ± 0.07</td>
<td>0.07 ± 0.08</td>
</tr>
<tr>
<td>calicheamicin</td>
<td>ds</td>
<td>0.038 ± 0.004</td>
<td>0.039 ± 0.004</td>
<td>0.007 ± 0.003</td>
<td>0.019 ± 0.009</td>
</tr>
</tbody>
</table>

* Model ds involves only single-strand cleavage, i.e., form \( I \) \( \rightarrow \) form II \( \rightarrow \) form III DNA. Model ds is for double-strand cleavage, form I \( \rightarrow \) form III and form III \( \rightarrow \) form I. 


a closed circular DNA as a carrier, it is evident that the affinity of DNase I for the closed circular DNA is much less than its affinity for linear DNA, e.g., calf thymus DNA. Since the enzyme must bind in order to cleave DNA, this suggests that $k_2'$ for cleavage of the closed circular form, will be lower than expected (due to lower affinity for closed circular DNA). The enzyme may also cause some double-strand cleavage of DNA, as has been shown to occur in the presence of MnCpO4. In the context of the single-strand cleavage model this would lead to a higher value of $k_2'$, thus also lowering the ratio $k_2'/k_2$. Although a model including both single- and double-strand cleavage could be used, it is unnecessary since a simple, single-strand cleavage model already fits the data within experimental error. It may be noted that $k_2'$ and $k_2$ are about the same size, suggesting that the process of conversion of form II is similar to the process of destruction of form III.

As shown in Table I and Figure 3, analysis of the data according to the double-strand cleavage model leads to a significantly poorer fit to the experimental intensities than using the single-strand model. This implies that the former model is not applicable and DNase I cleaves DNA primarily via a single-strand cleavage process. Further evidence for this conclusion is the appreciable negative (and thus unacceptable) value of $k_2'$ with the double-strand cutting model.

Esperamycin $A_1$. The single-strand cleavage model fits the data for this esperamycin analogue, as shown in Table I and Figure 4. An attempt to fit them with the double-strand model led to a 3-fold increase in the sum of the squared deviations, $D$, and to a substantial negative value of $k_2'$, showing that esp $A_1$ is mainly a single-strand cleavage agent. The ratio $k_2'/k_2$ for esp $A_1$ for the single-strand model is strikingly similar to that for DNase I (Table I). Because of the dramatic differences in structure and cleavage mechanism between the drug and the enzyme, this result was not expected. Like the enzyme, the drug may have a significantly reduced affinity for the tense form I DNA relative to the other two forms, which would lower the value of $k_2'$. Although the drug and the enzyme both cleave DNA from its minor groove, the drug is not likely to be as sensitive as the enzyme to the nuances of DNA structure. For example, changes in minor-groove width such as would occur between supercoiled form I and relaxed forms II and III DNA would be expected to affect drug binding to a lesser extent than enzyme binding. The fact that $k_2'/k_2$ is less than expected for random single-strand cleavage means that it also reflects a certain number of double-strand cleavages occurring with esp $A_1$. In the context of the single-strand model this would cause a higher value of $k_2'$ and lower the ratio $k_2'/k_2$. DNA cleavage by the drug, unlike the enzyme, is not expected to be catalytic. The activation mechanism requires the formation of a phenylene diradical that reacts with hydrogen atom donors in an irreversible manner, ultimately forming the aromatized drug esp Z. Esperamycin Z is not capable of cleaving DNA.

The single-strand cleavage behavior of esp $A_1$ means that one of the radicals of the phenylene diradical is unproductive with regard to a detectable strand break in the assay used. Although further study is warranted, it may be that the "silent" radicals react with solvent, the drug itself, or DNA. If the reaction with DNA involved the cleaved strand or the opposing strand but the hydrogen atom abstraction did not lead to a break in the sugar-phosphate backbone of the strand, form I would not convert directly to form III DNA. It may be noted that $k_2'$ is much bigger than $k_2$ for esp $A_1$. The reason for this is not understood, but our data show that, even at the highest esperamycin concentration, the concentration of form III DNA is still rising, which makes the determination of $k_2'$ from these data problematic.

Esperamycin C and Calicheamicin. Figures 5 and 6 and Table I show that the cleavage behavior of esp C and calicheamicin is significantly different from that of esp $A_1$ or DNase I. The sum of the squared deviations, $D$, in Table I shows that both models, single strand and double strand, fit the data for esp C and calicheamicin equally well. The fact that both models fit can be understood by considering the values of $k_2'$ and $k_2$ for the single-strand model. With $k_2' > k_2$ in this model, almost every initial break in a form I DNA species, converting it to form II, is quickly followed by the conversion of form II to form III. This makes the overall process mimicking double-strand or "single-strand" cleavage event: there is very little net production of form II because it is immediately converted to form III. This is indeed what is happening may be seen by comparing the calculated intensities for the two models (dashed lines and open symbols in Figures 5 and 6). The curves are practically identical, meaning that the single-strand model is equivalent to the double-strand model for this choice of parameters.

Obviously, one cannot conclude that double-strand cleavage is occurring to the exclusion of single-strand cleavage. Both probably take place, and one could construct a model including both processes I → II and I → III. Such a model would have an additional parameter (a rate constant). Since the present six-parameter models fit the data satisfactorily, fitting to a seven-parameter model does not seem a reasonable procedure. The value of $D$ would certainly decrease, but the values of the parameters determined would not be very meaningful.

In the double-strand cleavage model, the fact that $k_2'$ is much less than $k_2'$ implies that the formation of form III from form I is faster than its formation from form II. In the single-strand cleavage model, the ratio $k_2'/k_2$ reflects the ratio of the rate of II → III to the rate of I → II, large values of this ratio signify more effective double-strand cleavage. Since $k_2'/k_2$ = 3 for esperamicin C and 66 for calicheamicin, we conclude that the latter is more efficient at causing double-strand breaks. From the values of $k_2'$ it is also evident that the rate of cleavage of closed circular form I DNA by calicheamicin is 1 order of magnitude less than that of esperamicin C and almost 2 orders less than that of esperamicin $A_1$. Whether this is due to differences in the cleavage rate constants for the compounds or simply a manifestation of differing affinities that the agents may have for DNA remains to be determined. However, from the rate data it is clear that once calicheamicin cleaves one strand of DNA, cleavage of the opposing strand is rapid.

Although it is not possible to identify the relative locations of the two breaks in DNA from the agarose gel studies, earlier work using sequencing methods showed that double-strand induced breaks are three nucleotides apart and skewed in the 3'-direction along the strand. This indicates that the drug is bound in the minor groove and, as has been recently shown in modeling studies, with the warhead portion straddling two base pairs of DNA. Recent NMR studies have further indicated that the sugars of calicheamicin are "preorganized" and that they adopt a conformation outside of DNA closely matching that of the minor groove. Although less is known about the binding of esperamicin to DNA, the skewing in the drug's cleavage patterns and the ability of certain other drugs to influence DNA cleavage by esperamicin indicate that this compound also binds in the minor groove. The work presented here clearly shows that for esp C the warhead portion of the drug is situated such that targets on both strands are readily accessible and that double-strand cleavage occurs with high probability when compared to esp $A_1$. It is interesting to note that double-strand cleavage occurs with the two compounds having a single point of attachment between the oligosaccharide and the 1,5-dien-3-ene chromophore, e.g., esperamicin C and calicheamicin. Models (CPK) show that in the case of esperamicin $A_1$, which has substituents on two positions of the warhead, the deoxyribose and the attached anthranilate residue can be proximal to one of the radicals of the phenylene diradical. Although additional model building involving energy minimization will be nec-

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essary, this suggests that "self-quenching" by the drug of one of the radicals may be the reason that esp A₃ cleaves DNA in a single-strand manner.

Certain of the steps leading to the activation of the warhead portion of calicheamicin have recently been investigated by NMR methods.¹⁹ The first-order rate constant for the formation of aromatized drug from the Michael addition product outside of DNA is $5 \times 10^{-4}$ s⁻¹. In addition to being second-order rate constants, the $k''$ in Table 1 involve all of the individual steps including drug activation and the processes on DNA leading to strand breakage and thus cannot be directly compared with the rate constants of microsteps measured by NMR.

Esperamicin C and calicheamicin are unique in their DNA cleavage mechanisms. A single activation event on their chromophores leads to a double-strand break in DNA. This behavior is different from that of neocarzinostatin, which although it breaks one strand and modifies the second strand requires posttreatment of the DNA in order to produce a double-strand break.⁵ Since esp A₃ is more potent as an antitumor agent than is esp C,² single-versus double-strand cleavage is not the only factor influencing cytotoxicity. Cellular uptake and drug delivery very likely also play major roles in determining the biological effectiveness of the agents.

Conclusions

By using a coupled kinetic model, we studied the ability of esperamicin, calicheamicin, and the enzyme DNase I to cause single- or double-strand cleavage of DNA. The analysis shows that DNase I and esperamicin A₃ cleave DNA mainly via a single-strand process, whereas esperamicin C and calicheamicin cleave mainly through a double-strand mechanism. Examination of the structure of the drugs shows that the location of the sugars on the warhead portion of the agents is a factor influencing single-versus double-strand cleavage of DNA.

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Humidity-Controlled Reversible Structure Transition of Disodium Adenosine 5′-Triphosphate between Dihydrate and Trihydrate in a Single Crystal State

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Abstract: The humidity-controlled single-crystal transition of disodium adenosine 5′-triphosphate (Na₂ATP) between the dihydrate (1) and trihydrate (2) forms was investigated over the humidity range 5–50% at 23 °C by X-ray analysis. A crystal of 2 formed at high humidity changes into 1 at low humidity. The transition is reversible. The crystal structure of 1 was determined to be orthorhombic space group $P2_12_12_1$, with the cell parameters $a = 27.572 (5)$, $b = 21.066 (3)$, and $c = 7.0854 (9)$ Å. The structure was solved by direct methods and refined to the final $R$ value of 0.097 for 2450 independent reflections. Two water molecules, which are hydrogen-bonded to the hydroxyl groups of riboses in 2, are lost in 1. The triphosphate linkages are in helical arrangement, and the adenine bases are highly stacked along the $c$ axis, as for 2. There are two ATP molecules, A and B, in an asymmetric unit. The conformation of molecule A in 1 resembles that in 2; the ribose is C₃'-endo, and the exocyclic C₄'-C₅' torsion is gauche. In molecule B, the ribose is C₄'-endo, and the C₄'-C₅' torsion is gauche, in contrast to C₂'-endo-C₃'-exo, and gauche in 2. The torsional angles around the P-O ester bonds of the triphosphate chains of both molecules in 1 differ from those in 2 by 5–47°. The crystal structure of ATP makes the single-crystal transition possible. The conformation of molecule B in 1 relates to the conformation of ATP in its complexes with enzymes. The crystal structure of 2 (Kennard; et al. Proc. R. Soc. London 1971, A325, 401–436) was refined to the final $R$ value of 0.117 with the newly collected data, and the original assignment of Na₄ and OW₄ was interchanged.

Introduction

Adenosine 5′-triphosphate (ATP)-enzyme interactions have been discussed with regard to the crystal structures of enzyme-ATP (or ATP analogue) complexes.¹⁻² For example, in a series of site-directed mutagenic studies of tyrosyl-tRNA synthetase (TyrTS), the crystal structure of the TyrTS-tyrosyl adenylate complex was referred to in order to analyze the ATP-TyrTS interactions.⁵ Although the concept of a "rigid" nucleotide was proposed based on conformational analyses of the crystal structures of nucleotides,⁶ the adenosine moiety of ATP-enzyme complexes is sometimes in an "uncommon" conformation; the ribose of the